

THE FREEZING-POINT, BOILING-POINT, AND CONDUCTIVITY METHODS

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The Freezing-point, Boiling-point, and Conductivity Methods by Harry C. Jones

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HARRY C. JONES

**THE FREEZING-POINT,
BOILING-POINT, AND
CONDUCTIVITY METHODS**

©

... THE ...

Freezing-Point, Boiling-Point,

—AND—

Conductivity Methods

—BY—

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1897

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PREFACE

I have been impressed, in teaching the physical chemical methods in the laboratory, with the fact, that there is no readily accessible place in which they are treated satisfactorily from both the standpoint of theory and of practice. In the text-books, the theoretical side is developed, and usually without sufficient attention to the details of manipulation, to enable them to be applied successfully in the laboratory. In the laboratory manuals, on the other hand, these methods are often treated largely from the mechanical side, and their theoretical bearing might thus be lost sight of.

The physical chemical methods, which find most frequent application in the laboratory, are probably those based upon the lowering of the freezing-point, and the rise in the boiling-point of a solvent, produced by a dissolved substance, and the electrolytic conductivity of solutions of electrolytes. It is my chief object in preparing this little work to give an account of the operations involved in carrying out these methods in the laboratory. But since the mere mechanical application of any scientific method is a matter of comparatively little significance, I have aimed to give, also, enough of the theoretical ground on which each of them rests, to enable the student to work with them intelligently, and to see clearly their scientific significance and use.

HARRY C. JONES.

CONTENTS

PART I

THE FREEZING-POINT METHOD

	PAGE
Theoretical Discussion	I
Early History.....	1
Work of Raoult.....	1, 2
Molecular Lowering for Different Solvents.....	3
Molecular Lowering in Aqueous Solutions.....	4
Theory of Electrolytic Dissociation.....	5
Calculation of the Molecular Lowering.....	6, 7
Experimental Verification.....	8
Calculation of Molecular Weights from Lowering of Freezing-Point.....	8, 9
The Application of the Freezing-Point Method to the Determination of Molecular Weights in Solution	9
The Apparatus of Beckmann.....	10, 11
Carrying out a Determination.....	11-13
Correction for the Separation of Ice.....	13, 14
The Application of the Freezing-Point Method to the Measurement of Electrolytic Dissociation	14
The Method of Calculating Dissociation from Lowering of Freezing-Point.....	15, 16
The Method of Work.....	16
The Apparatus of Jones.....	17, 20
Comparison of the Results with the Dissociation from Conductivity Measurements.....	21

PART II

THE BOILING-POINT METHOD

	PAGE
Theoretical Discussion.....	23
Historical	23, 24
Work of Raoult.....	24, 25
The Relative Lowering of the Vapor-Tension	26
Calculation of Molecular Weights from Lowering of the Vapor-Tension	27
Beckmann's Work on Rise in Boiling-Point.....	27, 28
Calculation of Molecular Weights from Rise in the Boiling-Point of Solvents	28
Values of the Constants for Solvents.....	29
Relations between Boiling-Point and Freezing-Point Methods	29, 30
The Application of the Boiling-Point Method to the Determination of Molecular Weights in Solution.....	30
The Apparatus of Beckmann.....	31-33
The Apparatus of Hite	33-35
The Apparatus of Jones.....	34-36
Carrying Out a Determination.....	36-39
Correction for Separation of Vapor	39
Results of Measurements.....	40, 41

PART III

THE CONDUCTIVITY METHOD

Two Classes of Conductors.....	42
Electrolytes and Non-Electrolytes.....	42
Specific Conductivity	43, 44
Molecular Conductivity	44
Dissociation Measured by Conductivity Method.....	45, 46
Determination of μ_{∞}	46-50

CONTENTS

vii

	PAGE
The Application of the Conductivity Method to the Measurement of Electrolytic Dissociation	50
The Apparatus Employed	50-52
Calculation of the Molecular Conductivity.....	52, 53
Temperature Coefficient of Conductivity.....	54
The Ostwald Thermoregulator.....	55
Calibrating the Wire	56-58
Carrying Out a Conductivity Measurement	58
Determination of the Cell Constant.....	59
Precautions	60
Correction for the Conductivity of Water.....	60, 61
The Purification of Water	61-63
Substances to be Used.....	64

