

**POEMS. I. GRONGAR HILL.
II. THE RUINS
OF ROME. III. THE
FLEECE, IN FOUR BOOKS**

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JOHN DYER

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P O E M S.

BY

J O H N D Y E R, L. L. B.

V I Z.

I. G R O N G A R H I L L.

II. The R U I N S of R O M E.

III. The F L E E C E, in Four Books.



L O N D O N :

Printed for J. D O D S L E Y, in Pall-mall.

M. D C C. L X X.

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MR. JOHN DYER was born in Carmarthenshire, and educated at Westminster school. His father, an attorney of great practice and reputation, intended to introduce this his second son into his own business: but his genius led him a different way; besides his early taste for Poetry, he had a passion no less strong for the Arts of Design; and he determined to make Painting his profession.

With this view he made the voyage of Italy, where besides the usual study of the remains of Antiquity, and the works of the great Masters, he frequently spent whole days in the country about Rome and Florence, sketching those picturesque prospects with facility and spirit. Images from hence naturally transported themselves into his Poetical Compositions. The

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principal beauties of the *Ruins of Rome* are perhaps of this kind, and the various landscapes in the *Fleece* have been particularly admired.

On his return to England, he soon found, he could not relish a town life, nor submit to the assiduity required in his profession: his talent indeed was rather for Sketching than Finishing. So he contentedly sat down in the country with his little fortune, painting now and then a Portrait or a Landscape, as his fancy led him.

As his turn of mind was rather serious, and his conduct had been irreproachable, he very properly followed the opinion of his friends, who advised him to enter into Orders. And after some time spent upon a small cure in Warwickshire, his worthy character, and the merit of his Poetical Performances, recommended him to the notice of the Ld. Chancellor HARDWICKE, who presented him successively to the rectories of Belchford and Kerkby in Lincolnshire;

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colnshire; as did Sir JOHN HEATHCOTE to that of Coningsby in the same county. Upon this latter preferment he resided till the end of 1757, when a consumptive disorder, with which he had long struggled, carried him off at last in the 59th year of his age.

His character as a Writer has been freed by the following pieces published by himself, and now first collected: wherein a Poetical Imagination perfectly original, a Natural Simplicity, connected with and often productive of the True Sublime, and the warmest sentiments of Benevolence and Virtue, have been universally taken notice of. As a Member of Society, the same Simplicity appeared in his Manners, joined with a liberal turn of Thinking, which seldom solicited a Favor, and never lost a Friend.

The following table shows the results of the experiments conducted on the effect of temperature on the rate of reaction between hydrogen peroxide and potassium iodide. The reaction is catalyzed by the presence of a small amount of potassium iodide. The rate of reaction is measured by the volume of oxygen gas evolved in a given time.

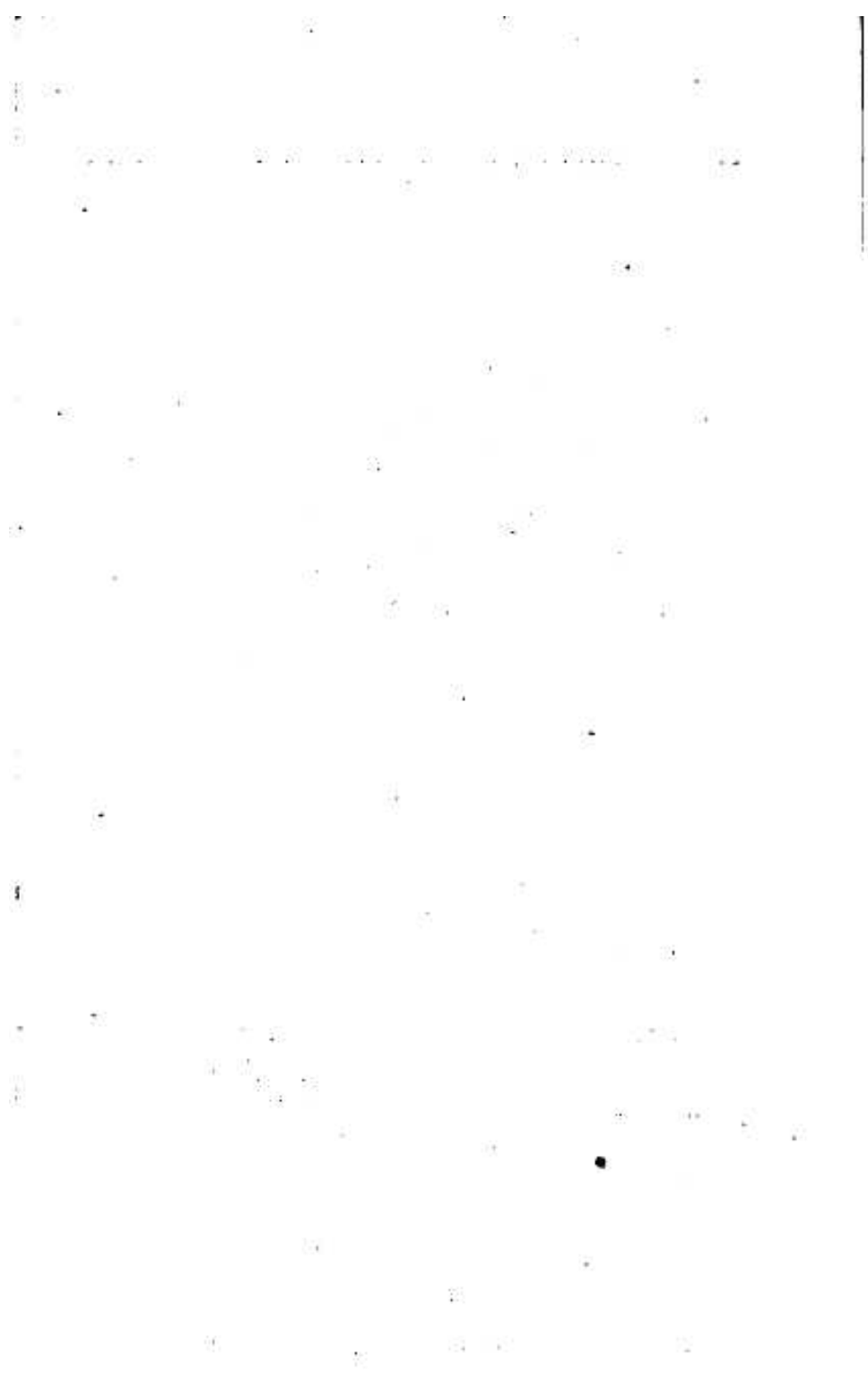
Temperature (°C)	Volume of O ₂ (cm ³)	Time (min)	Rate of Reaction (cm ³ min ⁻¹)
10	10	10	1.0
20	20	10	2.0
30	30	10	3.0
40	40	10	4.0
50	50	10	5.0
60	60	10	6.0
70	70	10	7.0
80	80	10	8.0
90	90	10	9.0

From the above table, it is clear that the rate of reaction increases with an increase in temperature. This is because the molecules of the reactants have more energy at higher temperatures, and therefore, they are more likely to collide with sufficient energy to overcome the activation energy barrier and undergo a chemical reaction.

The following table shows the results of the experiments conducted on the effect of concentration on the rate of reaction between hydrogen peroxide and potassium iodide. The reaction is catalyzed by the presence of a small amount of potassium iodide. The rate of reaction is measured by the volume of oxygen gas evolved in a given time.

Concentration of H ₂ O ₂ (M)	Volume of O ₂ (cm ³)	Time (min)	Rate of Reaction (cm ³ min ⁻¹)
0.1	10	10	1.0
0.2	20	10	2.0
0.3	30	10	3.0
0.4	40	10	4.0
0.5	50	10	5.0

From the above table, it is clear that the rate of reaction increases with an increase in the concentration of hydrogen peroxide. This is because there are more molecules of the reactants available for collision at higher concentrations, and therefore, the frequency of collisions is increased, leading to a higher rate of reaction.





Grongar Hill.